***Lesson Plans for the Week of: 12/12/16 Teacher: Hough Course: Chemistry Period: 9***

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| Elements of  a Lesson | **Monday** | **Tuesday** | **Wednesday** | **Thursday** | **Friday** |
| Objective/  Focus/  Essential  Question | CH.3a,d  --Name binary ionic compounds (with or without transition elements, especially without) | CH.2i; 3a,c,d  Quiz  Review for exam | CH.1a,b,h;2a,h,i;3a,c;4a;  5a,b,c,d,e  --Review for midterm exam: vocabulary/ concept practice | CH.1a,b,f,g,h;2a,h,i;3a,c;4a;  5a,b,c,d,e  Review for Midterm Exam | No Class |
| Lesson/Act.  Type of Presentation | Whole group:  Shortcut for yesterday’s lesson--connect charges of ions to the subscript in the molecular formula  Extension to yesterday’s lesson: Connect charges of ions to the columns of the periodic table using results of homework and #2 of U6 ws1  Notes: List properties of ionic compounds; illustrate conductivity of solutions using conductivity device  Note: explain the existence of a conductivity probe to quantitatively measure how well a substance conducts; this is opposite to the device I will use, which qualitatively determines whether or not a substance conducts | Quiz  Review for exam | Whole group:  Review vocabulary for thermal energy and categories of matter; include phase diagram and vapor pressure vocabulary, physical and chemical properties, changes  Student practice using concepts and vocabulary, particle diagrams | Individual:  Bellwork: Convert 8.1 mol CaCl2 to formula units; convert 4.3 g N2O to molecules  Whole Group:  Go over bellwork  Review Periodic Table items, then ions and ionic compounds: properties and naming |  |
| Evaluation | Teacher observation and results of student work |  | Teacher observation and results of student work | Teacher observation and results of student work |  |
| Extension/  Homework |  |  |  |  |  |

MATERIALS:

Monday: U6 ws1 from Chemistry Modeling Curriculum

Tuesday: Teacher-made mixed practice problems; teacher made topic review list

Wednesday: Teacher-made vocabulary worksheet; Practice differentiating between ionic and molecular compounds p. R54#61 all;

Write mixed formulas p.R54#66a,c-h; Write mixed names: p p.R54#65 all; Practice p. 299 #82#a-h

EXTRA: p. R54#67all—analyze why the item given cannot be solved as is; Potential Practice problems:

Calculate Molar Mass: p. R55#69a,b; #71a,b;#72b; convert Moles to Mass: p. R55 #75a,b,d; convert mass to moles: p. R56 #76a-c

convert moles to particles: p. R56 #77a-c; convert particles to moles: p. R56#78a-c

Thursday:

Friday: